

# Transition Pack for A Level Chemistry

**Get ready for A-level!**

**A guide to help you get ready for A-level Chemistry,  
including everything from topic guides to days out  
and online learning courses.**

**Commissioned by The PiXL Club Ltd. April 2018**

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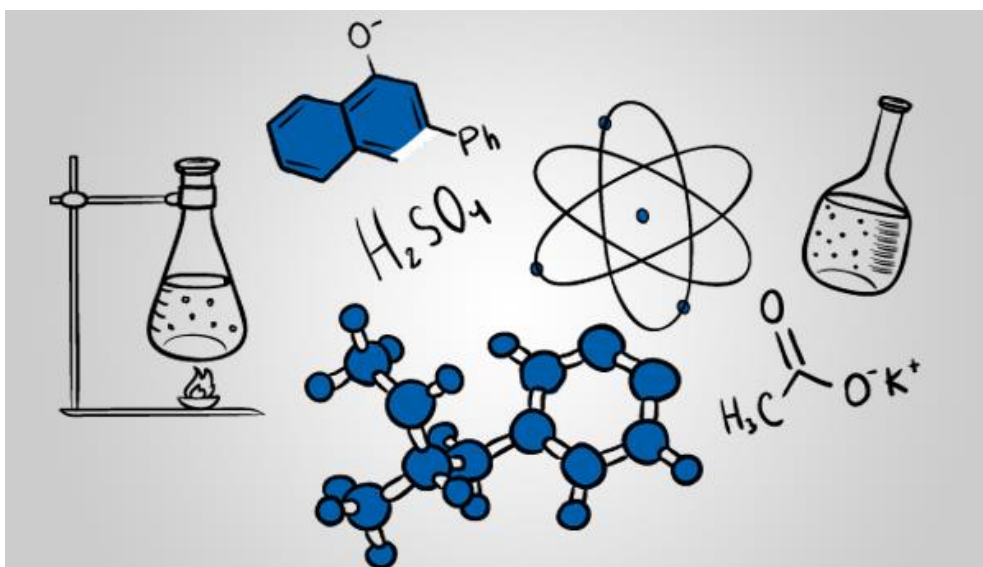
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[www.pixl.org.uk](http://www.pixl.org.uk)

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# So you are considering A level Chemistry?

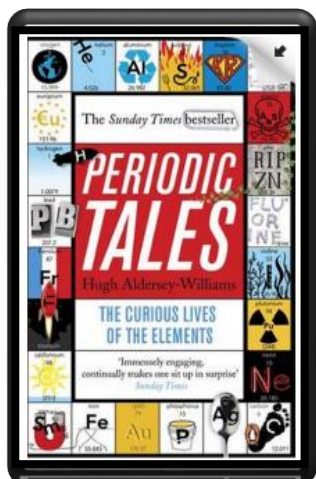
This pack contains a programme of activities and resources to prepare you to start A level in Chemistry in September. It is aimed to be used after you complete your GCSE, throughout the remainder of the summer term and over the summer holidays to ensure you are ready to start your course in September.



<https://www.my-mooc.com/en/mooc/chemistry1/>

## Book Recommendations

Kick back this summer with a good read. The books below are all popular science books and great for extending your understanding of chemistry

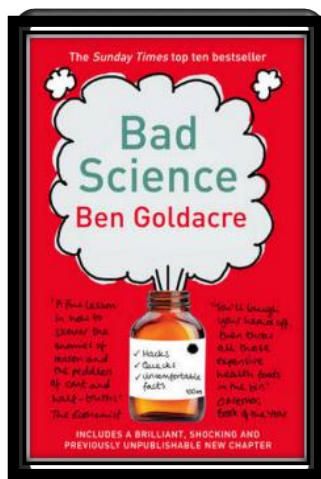
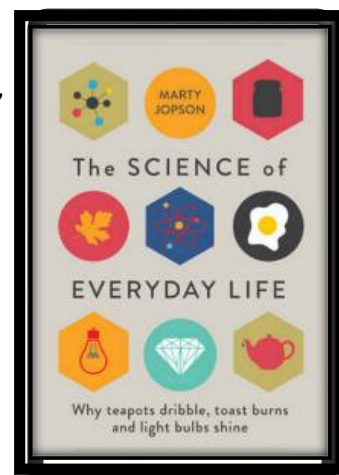


### Periodic Tales: The Curious Lives of the Elements

This book covers the chemical elements, where they come from and how they are used. There are loads of fascinating insights into uses for chemicals you would have never even thought about.

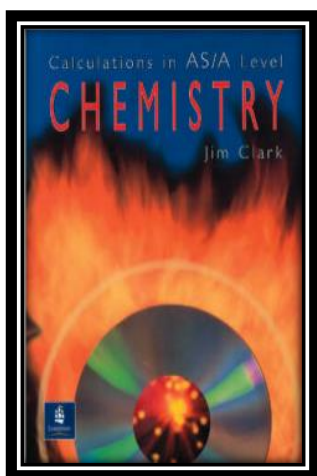
### The Science of Everyday Life: Why Teapots Dribble, Toast Burns and Light Bulbs Shine

The title says it all really, lots of interesting stuff about the things around your home!



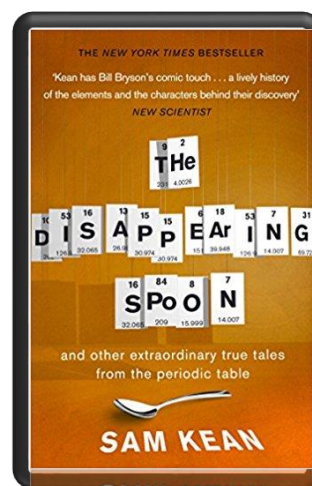
### Bad Science

Here Ben Goldacre takes apart anyone who published bad / misleading or dodgy science – this book will make you think about everything the advertising industry tries to sell you by making it sound 'sciencey'.



### Calculations in AS/A Level Chemistry

If you struggle with the calculations side of chemistry, this is the book for you. Covers all the possible calculations you are ever likely to come across. Brought to you by the same guy who wrote the excellent [chemguide.co.uk](http://chemguide.co.uk) website.

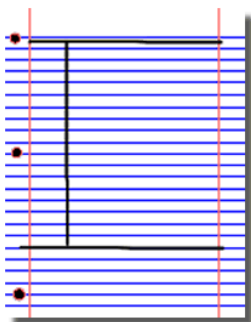


One of our crowning scientific achievements is also a treasure trove of passion, adventure, betrayal and obsession. **The Disappearing Spoon** follows the elements, their parts in human history, finance, mythology, conflict, the arts, medicine and the lives of the (frequently) mad scientists who discovered them.

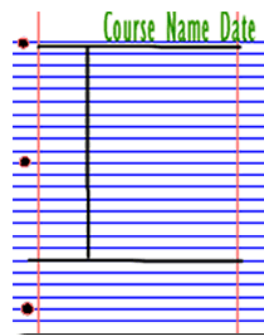
## Research Activities

Research, reading and note making are essential skills for A level chemistry study. For the following task you are going to produce 'Cornell Notes' to summarise your reading.

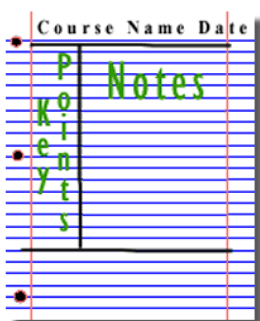
1. Divide your page into three sections like this



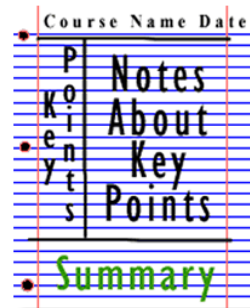
2. Write the name, date and topic at the top of the page



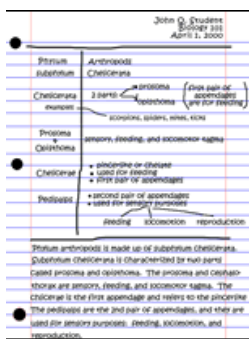
3. Use the large box to make notes. Leave a space between separate idea. Abbreviate where possible.



4. Review and identify the key points in the left hand box



5. Write a summary of the main ideas in the bottom space



Images taken from <http://coe.jmu.edu/learningtoolbox/cornellnotes.html>

## Research Activities

Aimed at students aged 14-19, Catalyst magazine is packed with interesting articles on cutting-edge science, interviews and new research written by leading academics. It also includes a booklet of teacher's notes, full of ideas and lesson plans to bring the articles to life in the classroom.

For each of the following topics you are going to use the resources to produce one page of Cornell style notes.

Use the links of scan the QR code to take you to the resources.

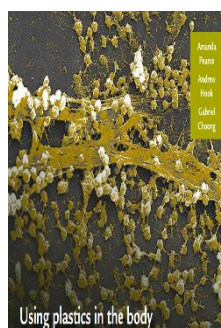


### Topic 1: Using Plastics in the Body

Available at:

<https://www.stem.org.uk/resources/elibrary/resource/382317/using-plastics-body>

This Catalyst article looks at how scientists are learning to use polymers for many medical applications, including implants, bone repairs and reduction in infections.



### Topic 2: Catching a Cheat

Available at:

<https://www.stem.org.uk/system/files/elibrary-resources/2017/03/Catching%20a%20cheat.pdf>

This Catalyst article looks at analytical chemists who are involved in many kinds of testing, including drug testing to catch cheats in sport.



### Topic 3: Diamond: More than just a gemstone

Available at:

<https://www.stem.org.uk/system/files/elibrary-resources/2017/02/Diamond%20more%20than%20just%20a%20gemstone.pdf>

This Catalyst article looks at diamond and graphite which are allotropes of carbon. Their properties, which depend on the bonding between the carbon atoms, are also examined.



## Chemistry Topic 1 – Electronic structure, how electrons are arranged around the nucleus

You will have used the rule of electrons shell filling, where:

7  
Li  
lithium  
3

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Elements that you expect to have a specific oxidation state actually have different states, so for example you would expect chlorine to be -1. It can have many oxidation states: NaClO, in this compound it has an oxidation state of +1

There are a few simple rules to remember:

Metals have a + oxidation state when they react.

Oxygen is 'king', it always has an oxidation state of -2.

Hydrogen has an oxidation state of +1 (except metal hydrides).

The charges in a molecule must cancel.

Examples: Sodium nitrate, NaNO<sub>3</sub>

Na +1      3x O<sup>2-</sup>  
+1      -6

sulfate ion, SO<sub>4</sub><sup>2-</sup>

4xO<sup>2-</sup> and 2- charges 'showing'  
-8      -2

To cancel: N = +5

S = +6

Q2. Work out the oxidation state of the underlined atom in the following:

- a) MgCO<sub>3</sub>      b) SO<sub>3</sub>      c) NaClO<sub>3</sub>      d) MnO<sub>2</sub>      e) Fe<sub>2</sub>O<sub>3</sub>      f) V<sub>2</sub>O<sub>5</sub>  
g) KMnO<sub>4</sub>      h) Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup>      i) Cl<sub>2</sub>O<sub>4</sub>

### Chemistry Topic 3 – Isotopes and mass

You will remember that isotopes are elements that have differing numbers of neutrons. Hydrogen has 3 isotopes; H<sub>1</sub><sup>1</sup>, H<sub>1</sub><sup>2</sup>, H<sub>1</sub><sup>3</sup>

Isotopes occur naturally, so in a sample of an element you will have a mixture of these isotopes. We can accurately measure the amount of an isotope using a **mass spectrometer**. You will need to understand what a mass spectrometer is and how it works at A level. You can read about a mass spectrometer here:

<http://bit.ly/pixlchem3>

<http://www.kore.co.uk/tutorial.htm>

<http://bit.ly/pixlchem4>

<http://filestore.aqa.org.uk/resources/chemistry/AQA-7404-7405-TN-MASS-SPECTROMETRY.PDF>

Q1. What must happen to the atoms before they are accelerated in the mass spectrometer?

Q2. Explain why the different isotopes travel at different speeds in a mass spectrometer.

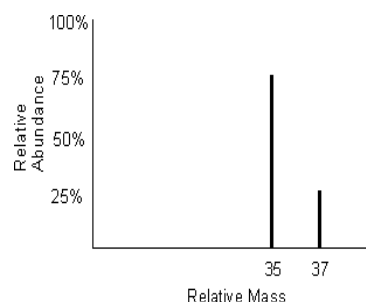
A mass spectrum for the element chlorine will give a spectrum like this:

75% of the sample consist of chlorine-35, and 25% of the sample is chlorine-37.

Given a sample of naturally occurring chlorine, ¾ of it

will be Cl-35 and ¼ of it is Cl-37. We can calculate what the **mean** mass of the sample will be:

$$\text{Mean mass} = \frac{75}{100} \times 35 + \frac{25}{100} \times 37 = 35.5$$



If you look at a periodic table, this is why chlorine has an atomic mass of 35.5.

An A level periodic table has the masses of elements recorded much more accurately than at GCSE. Most elements have isotopes and these have been recorded using mass spectrometers.

GCSE

11 <b>B</b> boron 5	12 <b>C</b> carbon 6	14 <b>N</b> nitrogen 7	16 <b>O</b> oxygen 8	19 <b>F</b> fluorine 9
27 <b>Al</b> aluminium 13	28 <b>Si</b> silicon 14	31 <b>P</b> phosphorus 15	32 <b>S</b> sulfur 16	35.5 <b>Cl</b> chlorine 17

A Level

10.8 <b>B</b> 5 boron	12.0 <b>C</b> 6 carbon	14.0 <b>N</b> 7 nitrogen	16.0 <b>O</b> 8 oxygen	19.0 <b>F</b> 9 fluorine
27.0 <b>Al</b> 13 aluminium	28.1 <b>Si</b> 14 silicon	31.0 <b>P</b> 15 phosphorus	32.1 <b>S</b> 16 sulphur	35.5 <b>Cl</b> 17 chlorine

Given the percentage of each isotope you can calculate the mean mass which is the accurate atomic mass for that element.

Q3. Use the percentages of each isotope to calculate the accurate atomic mass of the following elements.

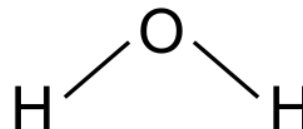
- Antimony has 2 isotopes: Sb-121 57.25% and Sb-123 42.75%
- Gallium has 2 isotopes: Ga-69 60.2% and Ga-71 39.8%
- Silver has 2 isotopes: Ag-107 51.35% and Ag-109 48.65%
- Thallium has 2 isotopes: Tl-203 29.5% and Tl-205 70.5%
- Strontium has 4 isotopes: Sr-84 0.56%, Sr-86 9.86%, Sr-87 7.02% and Sr-88 82.56%

#### **Chemistry Topic 4 – The shapes of molecules and bonding**

Have you ever wondered why your teacher drew a water molecule like this?

The lines represent a covalent bond, but why draw them at an unusual angle?

If you are unsure about covalent bonding, read about it here:



<http://bit.ly/pixlchem5>

<http://www.chemguide.co.uk/atoms/bonding/covalent.html#top>

At A level you are also expected to know how molecules have certain shapes and why they are the shape they are.

You can read about shapes of molecules here:

<http://bit.ly/pixlchem6>

<http://www.chemguide.co.uk/atoms/bonding/shapes.html#top>

Q1. Draw a dot and cross diagram to show the bonding in a molecule of aluminium chloride (AlCl<sub>3</sub>)

Q2. Draw a dot and cross diagram to show the bonding in a molecule of ammonia (NH<sub>3</sub>)

Q3. What is the shape and the bond angles in a molecule of methane (CH<sub>4</sub>)?

#### **Chemistry Topic 5 – Chemical equations**

Balancing chemical equations is the stepping stone to using equations to calculate masses in chemistry.

There are loads of websites that give ways of balancing equations and lots of exercises in balancing.

Some of the equations to balance may involve strange chemicals- don't worry about that, the key idea is to get balancing right.

<http://bit.ly/pixlchem7>

<http://www.chemteam.info/Equations/Balance-Equation.html>

This website has a download; it is safe to do so:

<http://bit.ly/pixlchem8>

<https://phet.colorado.edu/en/simulation/balancing-chemical-equations>

Q5. Balance the following equations

- $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- $\text{S}_8 + \text{O}_2 \rightarrow \text{SO}_3$
- $\text{HgO} \rightarrow \text{Hg} + \text{O}_2$
- $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
- $\text{C}_{10}\text{H}_{16} + \text{Cl}_2 \rightarrow \text{C} + \text{HCl}$
- $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
- $\text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- $\text{Fe}_2\text{O}_3 + \text{H}_2 \rightarrow \text{Fe} + \text{H}_2\text{O}$
- $\text{Al} + \text{FeO} \rightarrow \text{Al}_2\text{O}_3 + \text{Fe}$



### Chemistry Topic 6 – Measuring chemicals – the mole

From this point on you need to be using an A level periodic table, not a GCSE one. You can view one here:

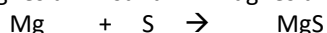
<http://bit.ly/pixlpertab>

[https://secondaryscience4all.files.wordpress.com/2014/08/filestore\\_aqa\\_org\\_uk\\_subjects\\_aqa-2420-w-trb-ptds\\_pdf.png](https://secondaryscience4all.files.wordpress.com/2014/08/filestore_aqa_org_uk_subjects_aqa-2420-w-trb-ptds_pdf.png)

Now that we have our chemical equations balanced, we need to be able to use them in order to work out masses of chemicals we need or we can produce.

The **mole** is the chemists equivalent of a dozen. Atoms are so small that we cannot count them out individually, we weigh out chemicals.

For example: magnesium + sulfur → magnesium sulfide



We can see that one atom of magnesium will react with one atom of sulfur. If we had to weigh out the atoms we need to know how heavy each atom is.

From the periodic table: Mg = 24.3 and S = 32.1

If I weigh out exactly 24.3g of magnesium this will be 1 mole of magnesium. If we counted how many atoms were present in this mass it would be a huge number ( $6.02 \times 10^{23}$ !!!!). If I weigh out 32.1g of sulfur then I would have 1 mole of sulfur atoms.

So 24.3g of Mg will react precisely with 32.1g of sulfur, and will make 56.4g of magnesium sulfide.

Here is a comprehensive page on measuring moles, there are a number of descriptions, videos and practice problems.

You will find the first 6 tutorials of most use here, and problem sets 1 to 3.

<http://bit.ly/pixlchem9>

<http://www.chemteam.info/Mole/Mole.html>

Q1. Answer the following questions on moles.

How many moles of phosphorus pentoxide ( $\text{P}_4\text{O}_{10}$ ) are in 85.2g?

How many moles of potassium are in 73.56g of potassium chlorate (V) ( $\text{KClO}_3$ )?

How many moles of water are in 249.6g of hydrated copper sulfate(VI) ( $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ )? For this one, you need to be aware the dot followed by  $5\text{H}_2\text{O}$  means that the molecule comes with 5 water molecules so these have to be counted in as part of the molecules mass.

What is the mass of 0.125 moles of tin sulfate ( $\text{SnSO}_4$ )?

If I have 2.4g of magnesium, how many g of oxygen ( $\text{O}_2$ ) will I need to react completely with the magnesium?  $2\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

### Chemistry Topic 7 – Solutions and concentrations

In chemistry a lot of the reactions we carry out involve mixing solutions rather than solids, gases or liquids.

You will have used bottles of acids in science that have labels saying 'Hydrochloric acid 1M', this is a solution of hydrochloric acid where 1 mole of HCl, hydrogen chloride (a gas) has been dissolved in  $1\text{dm}^3$  of water.

The  $\text{dm}^3$  is a cubic decimetre, it is actually 1 litre but from this point on as an A level chemist you will use the  $\text{dm}^3$  as your volume measurement.

<http://bit.ly/pixlchem10>

[http://www.docbrown.info/page04/4\\_73calcs11msc.htm](http://www.docbrown.info/page04/4_73calcs11msc.htm)

Q1.

- What is the concentration (in  $\text{mol dm}^{-3}$ ) of 9.53g of magnesium chloride ( $\text{MgCl}_2$ ) dissolved in  $100\text{cm}^3$  of water?
- What is the concentration (in  $\text{mol dm}^{-3}$ ) of 13.248g of lead nitrate ( $\text{Pb}(\text{NO}_3)_2$ ) dissolved in  $2\text{dm}^3$  of water?
- If I add  $100\text{cm}^3$  of  $1.00\text{ mol dm}^{-3}$  HCl to  $1.9\text{dm}^3$  of water, what is the molarity of the new solution?
- What mass of silver is present in  $100\text{cm}^3$  of  $1\text{mol dm}^{-3}$  silver nitrate ( $\text{AgNO}_3$ )?
- The Dead Sea, between Jordan and Israel, contains  $0.0526\text{ mol dm}^{-3}$  of Bromide ions ( $\text{Br}^-$ ). What mass of bromine is in  $1\text{dm}^3$  of Dead Sea water?

## Chemistry topic 8 – Titrations

One key skill in A level chemistry is the ability to carry out accurate titrations. You may well have carried out a titration at GCSE, at A level you will have to carry them out very precisely **and** be able to describe in detail how to carry out a titration - there will be questions on the exam paper about how to carry out practical procedures.

You can read about how to carry out a titration here, the next page in the series (page 5) describes how to work out the concentration of the unknown.

<http://bit.ly/pixlchem11>

[http://www.bbc.co.uk/schools/gcsebitesize/science/triple\\_aqa/further\\_analysis/analysing\\_substances/revision/4/](http://www.bbc.co.uk/schools/gcsebitesize/science/triple_aqa/further_analysis/analysing_substances/revision/4/)

Remember for any titration calculation you need to have a balanced symbol equation; this will tell you the ratio in which the chemicals react.

E.g. a titration of an unknown sample of sulfuric acid with sodium hydroxide.

A 25.00cm<sup>3</sup> sample of the unknown sulfuric acid was titrated with 0.100mol dm<sup>-3</sup> sodium hydroxide and required exactly 27.40cm<sup>3</sup> for neutralisation. What is the concentration of the sulfuric acid?

**Step 1:** the equation  $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$

**Step 2:** the ratios  $2 : 1$

**Step 3:** how many moles of sodium hydroxide  $27.40\text{cm}^3 = 0.0274\text{dm}^3$

number of moles =  $c \times v = 0.100 \times 0.0274 = 0.00274$  moles

**step 4:** using the ratio, how many moles of sulfuric acid

for every 2 NaOH there are 1 H<sub>2</sub>SO<sub>4</sub> so, we must have  $0.00274/2 = 0.00137$  moles of H<sub>2</sub>SO<sub>4</sub>

**Step 5:** calculate concentration. concentration = moles/volume  $\leftarrow$  in dm<sup>3</sup> =  $0.00137/0.025 = 0.0548 \text{ mol dm}^{-3}$

Here are some additional problems which are harder, ignore the questions about colour changes of indicators.

<http://bit.ly/pixlchem12>

<http://www.docbrown.info/page06/Mtestsnotes/ExtraVolCalcs1.htm>

Use the steps on the last page to help you.

Q1. A solution of barium nitrate will react with a solution of sodium sulfate to produce a precipitate of barium sulfate.

$\text{Ba}(\text{NO}_3)_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + 2\text{NaNO}_3(\text{aq})$

What volume of 0.25mol dm<sup>-3</sup> sodium sulfate solution would be needed to precipitate all of the barium from 12.5cm<sup>3</sup> of 0.15 mol dm<sup>-3</sup> barium nitrate?

### **Chemistry Topic 9 – Organic chemistry – functional groups**

At GCSE you would have come across **hydrocarbons** such as alkanes (ethane etc) and alkenes (ethene etc). You may have come across molecules such as alcohols and carboxylic acids. At A level you will learn about a wide range of molecules that have had atoms added to the carbon chain. These are called functional groups, they give the molecule certain physical and chemical properties that can make them incredibly useful to us.

Here you are going to meet a selection of the functional groups, learn a little about their properties and how we give them logical names.

You will find a menu for organic compounds here:

<http://bit.ly/pixlchem13>

<http://www.chemguide.co.uk/orgpropsmenu.html#top>

And how to name organic compounds here:

<http://bit.ly/pixlchem14>

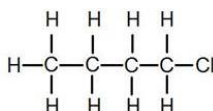
<http://www.chemguide.co.uk/basicorg/conventions/names.html#top>

Using the two links see if you can answer the following questions:

Q1. Halogenoalkanes

a. What is the name of this halogenoalkane?

b. How could you make it from butan-1-ol?



Q2. Alcohols

a. How could you make ethanol from ethene?

b. How does ethanol react with sodium and in what ways is this a) similar to the reaction with water, b) different to the reaction with water?

Q3. Aldehydes and ketones

a. Draw the structures of a) propanal, b) propanone

b. How are these two functional groups different?

### **Chemistry Topic 10 – Acids, bases, pH**

At GCSE you will know that an acid can dissolve in water to produce  $H^+$  ions, at A level you will need a greater understanding of what an acid or a base is.

Read the following page and answer the questions

<http://bit.ly/pixlchem15>

<http://www.chemguide.co.uk/physical/acidbaseeqia/theories.html#top>

Q1. What is your new definition of what an acid is?

Q2. How does ammonia ( $NH_3$ ) act as a base?

<http://bit.ly/pixlchem16>

<http://www.chemguide.co.uk/physical/acidbaseeqia/acids.html#top>

Q3 Ethanoic acid (vinegar) is a weak acid, what does this mean?

Q4 What is the pH of a solution of  $0.01 \text{ mol dm}^{-3}$  of the strong acid, hydrochloric acid?

## Science Websites

These websites all offer an amazing collection of resources that you should use again and again through out your course.

### chemguide

Helping you to understand Chemistry

MAIN MENU

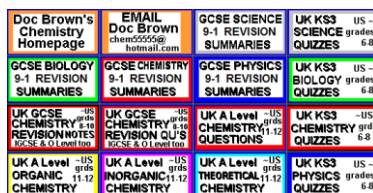
This website is very detailed and identifies other resources which are sharing incorrect or outdated information and suggests the correct materials to use. The site also contains links to the syllabuses of many exam boards which means it is accessible and useful to all students.

<https://www.chemguide.co.uk/>



The free revision website for students studying GCSE and A-levels. S-cool provides revision guides, question banks, revision timetable and more

<https://www.s-cool.co.uk/a-level/chemistry>



Doc Brown is a website dedicated to all three science subjects; physics, chemistry and biology. It provides the user with summarised notes (useful for making flash cards) and practice questions to further their knowledge and understanding.

The site provides resources from a wide range of exam boards including AQA, Edexcel, Chemistry, CCEA, OCR, WJEC, CIE and Salters from GCSE level to A2.

<http://www.docbrown.info/>

### chemrevise

Resources for A-level and GCSE Chemistry

HOME 1. AQA REVISION GUIDES 2. OCR REVISION GUIDES

5. A-LEVEL TEXTBOOK 6. GCSE AQA GUIDES ABOUT

Updates to A-level Textbook

The site was first made to host revision guides that are written for AQA A-level Chemistry. These revision guides have already been circulating on the internet for a couple of years on places like student room. This will be the place for the most up to date versions of them. The site has now extended to cover other exam boards. (OCR and Edexcel)

<https://chemrevise.org/>



Tons of awesome courses in one awesome channel! Check out the playlists for past courses in physics, philosophy, games, economics, U.S. government and politics, astronomy, anatomy & physiology, world history, biology, literature, ecology, chemistry, psychology, and of course, chemistry!

<https://www.youtube.com/user/crashcourse/featured>

## Pre-Knowledge Topics Answers to problems

Q1.

- |  |   |  |
|--|---|--|
| a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$         | b) $1s^2 2s^2 2p^6 3s^2 3p^1$                   | c) $1s^2 2s^2 2p^6 3s^2 3p^4$              |
| d) $1s^2 2s^2 2p^6 3s^2 3p^5$              | e) $1s^2 2s^2 2p^6 3s^2 3p^6$                   | f) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$    |
| g) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^3 4s^2$    | h) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$         | j) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^1$ |
| j) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$ | k) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$ |  |

Q2

- |                               |                                    |                                       |
|-------------------------------|------------------------------------|---------------------------------------|
| a) $1s^2 2s^2 2p^6 3s^2 3p^6$ | b) $1s^2 2s^2 2p^6 3s^2 3p^6$      | c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10}$ |
| d) $1s^2 2s^2 2p^6 3s^2 3p^6$ | e) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^7$ |                                       |

Q1

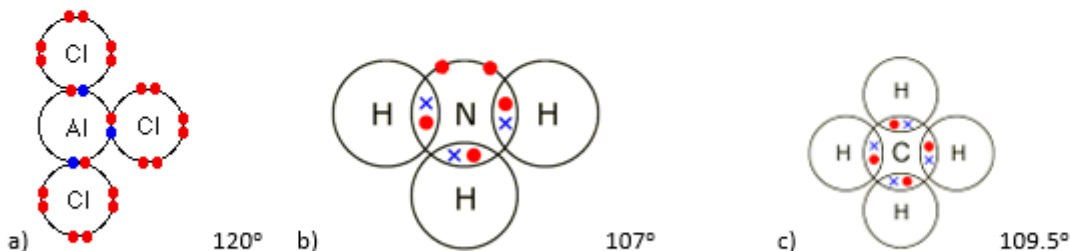
- a) +4   b) +6   c) +5   d) +4   e) +3   f) +5   g) +7   h) +6   j) +4

Q1 They must be ionised / turned into ions

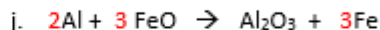
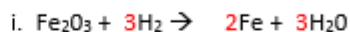
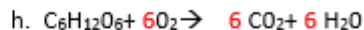
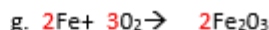
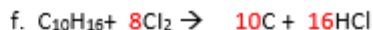
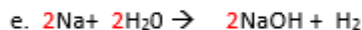
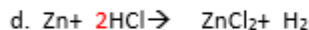
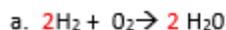
Q2 The ions are all given the same amount of kinetic energy, as  $KE = \frac{1}{2}mv^2$  the lighter ions will have greater speed / heavier ions will have less speed.]

- Q3   a) 121.855   b) 67.796   c) 107.973   d) 204.41   e) 87.710 / 87.7102

Q1



Q1



Q1

a)  $85.2/284 = 0.3 \text{ moles}$

b)  $73.56/122.6 = 0.6 \text{ moles}$

c)  $249.5/249.5 = 1.0 \text{ moles}$

d)  $0.125 \times 212.8 = 26.6\text{g}$

e)  $2\text{Mg} : 2\text{O}$  or 1:1 ratio     $2.4\text{g}$  of  $\text{Mg} = 0.1\text{moles}$     so we need  $0.1 \text{ moles}$  of oxygen ( $\text{O}_2$ ):  $0.1 \times 32 = 3.2\text{g}$

Q1

a)  $9.53\text{g}/95.3 = 0.1 \text{ moles}$ , in  $100\text{cm}^3$  or  $0.1\text{dm}^3$     in  $1\text{dm}^3$   $0.1\text{moles}/0.1\text{dm}^3 = 1.0 \text{ mol dm}^{-3}$

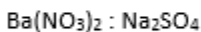
b)  $13.284\text{g}/331.2 = 0.04 \text{ moles}$ , in  $2\text{dm}^3$     in  $1\text{dm}^3$   $0.04\text{moles}/2\text{dm}^3 = 0.02 \text{ mol dm}^{-3}$

c)  $100\text{cm}^3$  of  $0.1 \text{ mol dm}^{-3} = 0.01 \text{ moles}$  added to a total volume of  $2 \text{ dm}^3 = 0.01\text{moles}/2\text{dm}^3 = 0.005 \text{ mol dm}^{-3}$

d) in  $1\text{dm}^3$  of  $1 \text{ mol dm}^{-3}$  silver nitrate,  $1 \text{ mole}$  of  $\text{Ag} = 107.9\text{g}$  in  $0.1\text{dm}^3 = 107.9 \times 0.1 = 10.79\text{g}$

e)  $0.0526 \times 79.7 = 42.0274\text{g}$

Q1



1 : 1 ratio

$12.5\text{cm}^3$  of  $\text{Ba}(\text{NO}_3)_2 = 0.0125\text{dm}^3$

$0.15 \text{ mol dm}^{-3} \times 0.0125\text{dm}^3 = 0.001875 \text{ moles}$

same number of moles of sodium sulfate needed, which has a concentration of  $0.25 \text{ mol dm}^{-3}$

$0.001875 \text{ moles} / 0.25 \text{ mol dm}^{-3} = 0.0075 \text{ dm}^3$  or  $7.5\text{cm}^3$



Q1 1-chlorobutane

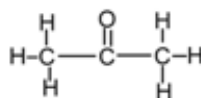
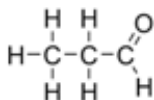
Add butan-1-ol to concentrated HCl and shake

Q2 React ethene with hydrogen gas at high temperature and pressure with a nickel catalyst

The reaction is similar in that it releases hydrogen but different as it proceeds much slower than in water

Q3 propanal

propanone

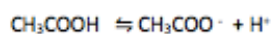


The carbon atom joined to oxygen in propanal has a hydrogen attached to it, it does not in propanone.

10.1 An acid is a proton donor

10.2 Ammonia can accept a proton, to become  $\text{NH}_4^+$

10.3 ethanoic acid has not fully dissociated, it has not released all of its hydrogen ions into the solution.



Mostly this      Very few of these

10.4  $\text{pH} = -\log [0.01] = 2$       The pH = 2



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